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Add a small amount of  
solute of known mass,  
if it dissolves it is  
unsaturated. If solid  
settles to the bottom  
filter the solution and  
mass the remaining  
solid, if the mass is the  
same then the...

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If solid settles to the bottom filter the solution and mass the remaining solid, if the mass is the same then the solution was saturated, if the mass is lower than it is an unsaturated solution.

Model 2 - Solute  
Dissolved vs. Solute  
Added The following  
data refer to an  
experiment in which a  
measured mass of solid  
is added to 10.0 g of  
20 °C water.

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## **2 Pogil Saturated And Unsaturated Solutions And Solubility ...**

solutions. This activity will explore whether or not there is a limit to how much of one substance can dissolve in another. Model

### 1 - Saturated and Unsaturated Solutions

- All beakers contain 10.0 g of water.
- All beakers are kept at 20 °C.
- All solutions are

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stirred for 2 hours. •  
Solute is the same  
substance in all  
beakers ...

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12 Identify the following variables in the experiment in Model 2 Dependent Lamar H S, Houston CHEMISTRY Chemistry - Fall 2014 ...

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unsaturated. If solid settles to the bottom filter the solution and mass the remaining solid, if the mass is the same then the solution was saturated, if the mass is lower than it is an unsaturated solution. Model 2 - Solute Dissolved

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bottom filter the  
solution and mass the  
remaining solid, if the  
mass is the same then



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## Unsaturated Saturated And Unsaturated Solutions Answers

### **Saturated And Unsaturated Solutions Pogil Answer Key**

Given scenarios, graphs, diagrams, or illustrations, the student will determine the type of solution such as saturated, supersaturated, or unsaturated.

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Saturated, Pogil  
Supersaturated, or**

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- A. Decreasing the temperature of the solution
- B. Increasing the temperature of the solution
- C. Removing water from the solution
- D. Stirring the solution

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There's more than one way to make a saturated solution. You can prepare it from scratch, saturate an unsaturated solution, or force a supersaturated solution to lose some solute. Add solute to a liquid until no more dissolves. Evaporate solvent from a solution until it becomes saturated.



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